

Catalysts for Reforming, Power to Liquids (PtL), Ammonia and Power to X pathways

Process	Reaction	Enthalpy of reaction	Temperature / °C	Typical catalysts	Commercial catalyst examples
Steam methane reforming (SMR)	$\text{CH}_4 + \text{H}_2\text{O} \rightarrow \text{CO} + 3\text{H}_2$	206 kJ/mol	700–950	NiO-based	Clariant REFORMAX 330 LDP
Oxygen-fed partial oxidation (POx)	$2\text{CH}_4 + \text{O}_2 \rightarrow 2\text{CO} + 4\text{H}_2$	– 24.5 kJ/mol	700–1100	CaAl ₁₂ O ₁₉	Clariant REFORMAX 420 with REFORMAX 330 LDP
High-temperature Water Gas Shift (HT WGS)	$\text{CO} + \text{H}_2\text{O} \rightarrow \text{CO}_2 + \text{H}_2$	– 131.2 kJ/mol	320–450	Fe ₂ O ₃ /Cr ₂ O ₃ with Cu promoter	Johnson Matthey KATALCO 71 range, Clariant ShiftMax 120
Low-temperature Water Gas Shift (LT WGS)	$\text{CO} + \text{H}_2\text{O} \rightarrow \text{CO}_2 + \text{H}_2$	– 131.2 kJ/mol	180–270	CuO & ZnO on alumina	Johnson Matthey KATALCO 83 range, Clariant ShiftMax 217
Reverse Water Gas Shift (rWGS)	$\text{CO}_2 + \text{H}_2 \rightarrow \text{CO} + \text{H}_2\text{O}$	41.5 kJ/mol	700–900	Ni-based	Clariant HyProGen R-70
Catalytic POx rWGS	$2\text{CO}_2 + 2\text{O}_2 + 6\text{H}_2 \rightarrow 2\text{CO} + 6\text{H}_2\text{O}$	$\Delta H < 0$	800–950	Ni on metal oxide, Rh on alumina	Johnson Matthey HyCOgen, VTT
Dry Methane Reforming (DMR)	$\text{CO}_2 + \text{CH}_4 \rightarrow 2\text{CO} + 2\text{H}_2$	247.1 kJ/mol	800–950	Ni oxides	BASF SYNSPIRE G1-110, Clariant HyProGen R-70
Methanation of CO ₂	$\text{CO}_2 + 4\text{H}_2 \rightarrow \text{CH}_4 + 2\text{H}_2\text{O}$	– 165.0 kJ/mol	300–400	NiO/alumina	Clariant METH 135, Topsøe MCR-2
Methanol Synthesis from Syngas	$\text{CO} + 2\text{H}_2 \rightarrow \text{CH}_3\text{OH}$	– 90.4 kJ/mol	200–300	CuO/ZnO/MgO on alumina	Topsøe MK-121, Johnson Matthey KATALCO 51, Clariant MegaMax 900
Methanol Synthesis via CO ₂ -Hydrogenation	$\text{CO}_2 + 3\text{H}_2 \rightarrow \text{CH}_3\text{OH} + \text{H}_2\text{O}$	– 49.5 kJ/mol	200–300	Cu-based	Johnson Matthey eMERALD, Clariant MegaMax 800
Methanol to Olefins (ethylene)	$2\text{CH}_3\text{OH} \rightarrow \text{C}_2\text{H}_4 + 2\text{H}_2\text{O}$	– 11.72 kJ/mol	400–500	SAPO-34, ZSM-5	UOP MTO-600
Methanol to Olefins (propylene)	$3\text{CH}_3\text{OH} \rightarrow \text{C}_3\text{H}_6 + 3\text{H}_2\text{O}$	– 30.98 kJ/mol	400–500	SAPO-34, ZSM-22	Clariant MTPROP
Methanol reforming to syngas	$\text{CH}_3\text{OH} + \text{H}_2\text{O} \rightarrow 3\text{H}_2 + \text{CO}_2$	49,2 kJ/mol	240–320	CuO/ZnO/MgO on alumina	Topsøe MDK-20
DME Synthesis from methanol	$2\text{CH}_3\text{OH} \rightarrow \text{CH}_3\text{OCH}_3 + \text{H}_2\text{O}$	– 23.5 kJ/mol	250–300	Al ₂ O ₃ or zeolite based	Topsøe DME-99-ECO
DME to Paraffins	$\text{CH}_3\text{OCH}_3 \rightarrow 2(-\text{CH}_2-) + \text{H}_2\text{O}$	$\Delta H < 0$	320–380	Zeolite based, with iron promoter	Nanostructured ZSM-5
Formic Acid Synthesis via CO ₂ -Hydrogenation	$\text{CO}_2 + \text{H}_2 \rightarrow \text{HCOOH}$	– 31.2 kJ/mol	<100	RuCl ₂ (PTA) ₄ , CU-MOF-5	Emerging
CO-Based Fischer-Tropsch Synthesis (CO FTS)	$\text{CO} + 2\text{H}_2 \rightarrow (-\text{CH}_2-) + \text{H}_2\text{O}$	– 152 kJ/mol	180–250	Co-, Fe-based	Johnson Matthey FT CANS, Greyrock GreyCat
CO ₂ -Based Fischer-Tropsch Synthesis (CO ₂ FTS)	$\text{CO}_2 + 3\text{H}_2 \rightarrow (-\text{CH}_2-) + 2\text{H}_2\text{O}$	– 111 kJ/mol	300–350	Na or K modified Co-Fe based	Emerging
Ammonia synthesis (Haber Bosch)	$\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$	– 45,9 kJ/mol NH ₃	350–550	Fe-based	Topsøe KM1
Low temperature ammonia cracking	$2\text{NH}_3 \rightarrow \text{N}_2 + 3\text{H}_2$	45,9 kJ/mol NH ₃	400–500	Ru, PGM-based	Johnson Matthey KATALCO 27-612, Clariant HyProGen 850 DCARB
High temperature ammonia cracking	$2\text{NH}_3 \rightarrow \text{N}_2 + 3\text{H}_2$	45,9 kJ/mol NH ₃	600–800	Ni-based	Johnson Matthey KATALCO 27-2, Topsøe H ₂ Retake, Clariant HyProGen 820 DCARB
PEM electrolyser cathode reaction (HER)	$4\text{H}^+ + 4\text{e}^- \rightarrow 2\text{H}_2$	285.83 kJ/mol water	50–80	Pt	Pajarito Powder Pt/ECS, Johnson Matthey Platinum black HSA
PEM electrolyser anode reaction (OER)	$2\text{H}_2\text{O} \rightarrow 4\text{H}^+ + 4\text{e}^- + \text{O}_2$	285.83 kJ/mol water	50–80	Ir	Pajarito Powder EEC-IrOx, Johnson Matthey Iridium-black
PEM fuel cell cathode reaction	$4\text{H}^+ + 4\text{e}^- + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$	– 285.83 kJ/mol water	ambient–55	Pt-based	Pajarito Powder Pt/ECS-3701, Heraeus H ₂ FC-50Pt-C700
PEM fuel cell anode reaction	$\text{H}_2 \rightarrow 2\text{H}^+ + 2\text{e}^-$	– 285.83 kJ/mol water	ambient–55	Pt	Pajarito Powder Pt/ECS, Heraeus H ₂ FC-30Pt-C60T
AEM electrolyser cathode reaction (HER)	$2\text{H}_2\text{O} + 2\text{e}^- \rightarrow 2\text{OH}^- + \text{H}_2$	285.83 kJ/mol water	ambient–55	non-PGM	Pajarito Powder Pt/ECS
AEM electrolyser anode reaction (OER)	$4\text{OH}^- \rightarrow 2\text{H}_2\text{O} + \text{O}_2 + 4\text{e}^-$	285.83 kJ/mol water	ambient–55	Ni, Co, Fe-based	Pajarito Powder EEC-PbRuOx